Chemistry 115 Name

Dr. Cary Willard

Quiz 6a (20 points) April 2, 2014

1. (6 points) Write and balance the double displacement reaction that occurs between sodium phosphate and magnesium chloride. Remember that phosphates are generally insoluble and chlorides are generally soluble.

Na3PO4(aq) + MgCl2(aq) 🡪

1. (4 points) How does an orbit differ from an orbital?
2. (2 points) How many electrons can occupy a 6p orbital?
3. (4 points) What is the complete electron configuration of an atom of silicon?
4. (4 points) What is the shorthand electron configuration of an atom of cadmium (Cd)?

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Quiz 6b (20 points) April 2, 2014

1. (6 points) Write and balance the double displacement reaction that occurs between potassium hydroxide and aluminum chloride. Remember that hydroxides are generally insoluble and chlorides are generally soluble.

KOH(aq) + AlCl3(aq) 🡪

1. (4 points) How does an orbit differ from an orbital?
2. (2 points) How many electrons can occupy a 5f orbital?
3. (4 points) What is the complete electron configuration of an atom of sulfur?
4. (4 points) What is the shorthand electron configuration of an atom of iron (Fe)?

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Quiz 6c (20 points) April 2, 2014

1. (6 points) Write and balance the double displacement reaction that occurs between lithium carbonate and iron(III) bromide. Remember that carbonates are generally insoluble and bromides are generally soluble.

Li2CO3(aq) + FeBr3(aq) 🡪

1. (4 points) How does an orbit differ from an orbital?
2. (2 points) How many electrons can occupy a 7d orbital?
3. (4 points) What is the complete electron configuration of an atom of phosphorus?
4. (4 points) What is the shorthand electron configuration of an atom of zirconium (Zr)?

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Quiz 6d (20 points) April 2, 2014

1. (6 points) Write and balance the double displacement reaction that occurs between potassium phosphate and barium iodide. Remember that phosphates are generally insoluble and iodides are generally soluble.

K3PO4(aq) + BaI2(aq) 🡪

1. (4 points) How does an orbit differ from an orbital?
2. (2 points) How many electrons can occupy a 3p orbital?
3. (4 points) What is the complete electron configuration of an atom of chlorine?
4. (4 points) What is the shorthand electron configuration of an atom of vanadium (V)?